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well as the chemical reactions, interferences, and calculations will be studied. There is a major emphasis on chemical equilibrium as a basis for understanding the design of analytical methods, and also on the application of chemical knowledge and stoichiometric relationships to the solution of problems in chemical analysis.

CHEM 2222 - Quantitative Analysis

3 LE CHATELIER'S PRINCIPLE $\text{NaCl(s)} + \text{H}_2\text{SO}_4\text{(l)} \rightarrow \text{Na}_2\text{SO}_4\text{(s)} + \text{HCl(g)}$ reaction is carried out in an open container Because HCl is a gas that can escape from the system,

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